

## CHEMISTRY 112, INTRODUCTORY CHEMISTRY

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### COURSE MATERIALS

1. Text: Chemistry: Structure and Dynamics, 4th Edition; J. N. Spencer, G. M. Bodner and L. H. Rickard; John Wiley & Sons, 2008, ISBN-13 978-0-470-129289
2. Laboratory Manual: Thomas G. Greco, Lyman H. Rickard and Gerald S. Weiss, Experiments in General Chemistry, Principles and Modern Applications, 9th ed., Prentice Hall, Upper Saddle River, NJ, 2007, ISBN. 0-13-149391-4.
3. Chemistry: A Guided Inquiry, 4th Edition; Richard S. Moog and John J. Farrell, John Wiley & Sons, 2008.
4. Laboratory Notebook: Permanent bound (no ring binders), quadrille-ruled, approximately 7 x 9.5".
5. Safety Goggles/Glasses & lock for your lab drawer: Available from the ACS Student Affiliates in the General Chemistry Prep-room (Caputo 330) or from the bookstore.

A C- or better in CHEM 111 is a prerequisite for this class.

### COURSE OUTLINE

Chapter	Topic
10	Kinetics and Equilibrium Reactions That Don't Go to Completion Gas-Phase Reactions The Rate of a Chemical Reaction The Collision Theory of Gas-Phase Reactions Equilibrium Constant Expressions Reaction Quotients Changes in Concentration That Occur as a Reaction Comes to Equilibrium The Effect of Temperature on an Equilibrium Constant LeChâtelier's Principle Equilibrium Reactions Which Involve Pure Solids and Liquids
11	Acids and Bases Properties of Acids and Bases The Arrhenius & Bronsted Definitions of Acids and Bases Conjugate Acid--Base Pairs The Role of Water in the Bronsted Model pH as a Measure of the Concentration of the $\text{H}_3\text{O}^+$ Ion The Relative Strengths of Acids and Bases Relationship of Structure to Relative Strengths of Acids Acid and Base pH Calculations Buffers and Buffer Capacity
12	Oxidation-Reduction Oxidation Numbers Recognizing Oxidation--Reduction Reactions Voltaic Cells Oxidizing and Reducing Agents Relative Strengths of Oxidizing and Reducing Agents Standard-State Cell Potentials Nonstandard Conditions & The Nernst Equation Electrolysis and Faraday's Law
13	Thermodynamics Spontaneous Chemical and Physical Processes Entropy as a Measure of Disorder Entropy and the Second Law of Thermodynamics The Third Law of Thermodynamics Calculating Entropy Changes for Chemical Reactions Gibbs Free Energy The Effect of Temperature on the Free Energy of a Reaction Equilibria Expressed in Partial Pressures Interpreting Standard-State Free Energy of Reaction Data

- The Relationship Between Free Energy and Equilibrium Constants
- The Temperature Dependence of Equilibrium Constants
- 14 Rates of Chemical Reaction
  - The Forces That Control a Chemical Reaction
  - Chemical Kinetics
  - Instantaneous Rates of Reaction
  - Rate Laws and Rate Constants
  - Order and Molecularity
  - Collision Theory of Chemical Reactions
  - The Mechanisms of Chemical Reactions
  - Determining the Order of a Reaction
  - The Integrated Form of First-Order and Second-Order Rate Laws
  - The Activation Energy of Chemical Reactions
  - Catalysts and the Rates of Chemical Reactions

### TENTATIVE EXAM SCHEDULE

- Exam 1: Chapters 10 & 11: Oct 2
- Exam 2: Chapters 12 & 13: Nov 6
- Exam 3: Chapter 14: Dec 2
- Final Exam: Cumulative: Dec 10 at 2:45-4:45

### LABORATORY SCHEDULE

Date	Expt #	Title
Aug 26, 27	16	Determination of an Equilibrium Constant, Part A
Sept 2, 3	16	Determination of an Equilibrium Constant, Part B
Sept 9, 10	14	Equilibrium: Le Chatelier's Principle
Sept 16, 17	19	Determination of the Ionization Constant of a Weak Acid
Sept 23, 24	18	Weak Acids, Bases, and Their Salts, Method III
Sept 30, Oct 1	20	Investigation of Buffer Systems
Oct 8, 14	33-35	Qual I
Oct 15, 21	28	A Penny's Worth of Chemistry
Oct 22, 28	25	Investigation of Voltaic Cells, Parts A & B
Oct 29, Nov 4	33-35	Qual I
Nov 5, 11	33-35	Qual I
Nov 12, 18	Handout	Crystal Violet Kinetics
Nov 19, 25	15	Kinetics Study of an Iodine Clock Reaction
Dec 2, 3	32	Paper Chromatography, Parts A & C

### HOMEWORK ASSIGNMENTS

Read each chapter as it is covered in lecture.

Chapter 10: 1, 3, 11, 12, 13, 14, 17, 19, 23, 25, 27, 29, 35, 39, 41, 49, 50, 53, 54, 58, 61, 63, 65, 73, 75, 77, 81, 85, 114, 117, 119

Chapter 11: 3, 7, 13, 15, 17, 19, 21, 35, 37, 41, 48, 51, 54, 55, 63, 64, 67, 68, 69, 70, 77, 83, 85, 87, 89, 95, 97, 98, 105, 106, 107, 108, 111, 113, 115, 117, 121, 125, 141, 148,

Chapter 12: 1, 3, 5, 9, 13, 15, 31, 33, 43, 53, 55 (omit c), 56, 58, 60, 61, 62, 63, 64, 65, 69, 82, 85, 86, 91, 93, 98, 105, 131

Chapter 13: 1, 9, 11, 15, 21, 23, 28, 29, 31, 33, 39, 43, 45, 49, 50, 51, 55, 59, 63, 72, 73, 79, 83,

Chapter 14: 1, 12, 13, 15, 16, 21, 23, 26, 29, 30, 31, 33, 35, 37, 44, 45, 49, 51, 52, 53, 55, 59, 63, 65, 85, 86, 87, 88, 89, 91, 92, 93

### GRADING SYSTEM

Exams (3)	300 pts
Quizzes\Worksheets	100 pts (The lowest quiz grade will be dropped.)
Final Exam	200 pts
Laboratory	<u>200 pts</u>
Total	800 pts

Letter Grade

Points

Percentage

A	736 - 800	92 - 100
A-	720 - 735	90 - 92
B+	704 - 719	88 - 90
B	656 - 703	82 - 88
B-	640 - 655	80 - 82
C+	616 - 639	77 - 80
C	544 - 615	68 - 77
C-	520 - 543	65 - 68
D+	512 - 519	64 - 65
D	488 - 511	61 - 64
D-	480 - 487	60 - 61
F	Below 480	< 60

A passing grade in the lecture component (360 pts of the possible 600) of the course must be achieved in order to receive a passing grade.

### **OFFICE HOURS**

Monday: 3:00-4:00; Tuesday and Thursday: 11:00-12:00 & 2:00-3:00. Feel free to come by any time that I am in my office. If you have difficulty finding a time to meet with me, make an appointment to see me.

### **ATTENDANCE**

Regular attendance is expected for all lectures, recitations and laboratories. If an absence results in a missed exam, quiz or lab, a request for excused absence form (attached) must be submitted to the instructor. Absences are excused for illness, family emergencies or university activities. Absences due to university activities must be discussed with the instructor in advance and arrangements made for making up the missed work. If an exam or quiz is missed and the absence is excused, the final exam grade will be substituted for the missed work.

### **STUDENT RESPONSIBILITIES**

You are responsible for all assigned work and material covered in class and lab. Work submitted after the due date will receive a grade of F unless prior arrangements have been made. However, the work must still be submitted or a grade of zero will be recorded. The average student will need to spend approximately 1 hour and 15 minutes preparation time for the course every day (including weekends) outside of class in order to receive an average grade. A higher grade will normally require additional study time. Two days of 5 hours each are not equivalent to an hour and 15 minutes each day. You should diligently prepare all assignments. When you encounter material which you can not master alone you should seek help immediately. The primary source of help should be the course instructor. You will find my office hours listed in the syllabus and posted on my office door. If you are having difficulty in the course I expect you to come see me.

### **PROBLEM ASSIGNMENTS/QUIZZES**

Frequent problem assignments will be made from the text. The problem assignments will not be collected and graded. Instead there will be frequent quizzes in lecture or recitation taken from the assigned problems and recitation worksheets. You are expected to read each chapter as it is being covered in lecture. Study the chapter in detail to increase your understanding of the material. In some cases reading assignments will be made which will not be covered in lecture. Study all assigned homework problems until you understand them (not just until you obtain an answer). Plan to work only a few new problems (5-8) each night plus reading (60 minutes). Then review the problems from the previous night (15 minutes). You should be confident of the homework assignment before coming to class. The homework problems assigned are the minimum number of problems you are expected to work. They represent a broad overview of the types of problems you are expected to understand. You are expected to work additional problems and to spend additional time on those problems that give you difficulty.

### **LABORATORY**

You should read the lab safety rules found in the introduction of the laboratory manual. A copy of these rules must be signed and returned to the instructor at the first lab. You should bring your laboratory manual, laboratory notebook, and safety goggles to all labs. Laboratory experiments should be done during the assigned laboratory period. Any work outside of this period must be approved by the lab instructor; a chemistry faculty member must be available to supervise your work; and another student must be in the lab with you at all times. Before coming to lab you are expected to have diligently studied the experiment and outlined the procedure in your laboratory notebook. The pre-lab assignment should be completed but will not be handed in to the instructor. There will be a pre-lab quiz given in lab at the beginning of

each new experiment.

### **CLASSROOM ETIQUETTE**

Arrive for class, recitation and laboratory on time. Turn off all cell phones while in class, recitation or lab. Cell phones may *not* be used as a calculator during quizzes or tests. Talking during lecture is a distraction to others who are trying to listen.

### **ACADEMIC HONESTY**

Students are expected to be familiar with the University's policy on academic honesty and dishonesty found in the *Student Handbook* and the *Academic Honesty and Dishonesty* brochure. Academic dishonesty includes cheating on an exam or quiz, presenting another student's work as one's own in a laboratory report or notebook, fabricating data in a laboratory experiment. Students are encouraged to work together on homework assignments, preparation for exams and laboratory reports. However, the results submitted in a laboratory report must represent the students own work.

### **THE LABORATORY NOTEBOOK**

The laboratory notebook is a permanent record of your work in the laboratory. You must have your notebook with you in order to work in the lab. All notebooks must be permanently bound and begin with a table of contents. All entries should be in ink. Each page must be consecutively numbered. As each page is completed you should initial and date the bottom of the page. The procedure must be outlined in the notebook before coming to lab. All data must be recorded in the notebook using correct significant figures and proper units. Never write data on another sheet of paper with the idea of transferring it to the notebook. Notebooks should be relatively neat and orderly, however, data should never be recopied into another notebook. If an error is made, do not obliterate the data (also do not use white out, tear out pages or tape in new pages). Draw a single line through any errors and write the correct data.

The notebook is a record of your work as it is done. The notebook should be kept in such a way that the instructor can turn to any experiment and tell exactly what you did during the experiment. All data must include the appropriate units and be labeled to identify the data. All calculations, graphs, tables and assigned questions must be included in the notebook. You should review pages ten and eleven of your laboratory manual for more information on the laboratory notebook. All lab notebooks should be reviewed and initialed by the instructor at the end of each laboratory period.

A conclusion will be required for some lab reports. The conclusion should be one paragraph. It should state the major results of the experiment. This statement of results should agree with the purpose of the experiment written at the beginning of the report. This should be followed by a statement describing whether you are confident in the results. The remainder of the conclusion should be an argument to convince the reader why you feel your results are appropriate or not. This argument can refer to the agreement between multiple trials, agreement with other student results, trends in the data such as a linear graph or other observations from the experiment.

Laboratory notebooks will be evaluated on the following criteria:

1. Top of each page: title and page #; Bottom of each page: date completed and initials
2. Format: organization, neatness, completeness
3. Purpose: describe what is being measured or determined
4. Data Presentation: significant figures, labels and units
5. Results: quality of results
6. Conclusions: one paragraph (when requested)
7. Questions
8. Graphs: title, labels, units

## Request for Excused Absence

Student Name:

Dates of Absence:

Reason for Absence (circle one): Illness, Family Emergency, University Activity

I request this absence be excused and that: (check all that apply)

\_\_\_\_\_ my final exam grade be substituted for the missed lecture quiz.

\_\_\_\_\_ my final exam grade be substituted for the missed test.

\_\_\_\_\_ my final exam grade be substituted for the missed pre-lab quiz.

\_\_\_\_\_ I be allowed to make-up the missed laboratory experiment. The lab must be made-up and the laboratory report submitted for grading within one week of returning to class.

Attach documentation to support the request for an excused absence.